

**THE UNITED REPUBLIC OF TANZANIA  
NATIONAL EXAMINATIONS COUNCIL OF TANZANIA**

**032 CHEMISTRY ACTUAL PRACTICALS FOR ORDINARY LEVEL  
OF SECONDARY EDUCATION**

**1.0 IMPORTANT**

The National Examinations Council has prepared a checklist of apparatus and chemicals for Chemistry Actual Practicals. As a Head of the school **make sure that all the apparatus and chemicals indicated in this checklist are available** in the school laboratory. Some of these will be used for Certificate of Secondary Education Examination (CSEE) 2018 Chemistry practicals. There **will be no one month** Advance Instructions which will be provided for Chemistry practicals. However, 24 Hours Advance Instructions will be provided.

**2.0 LIST OF APPARATI AND CHEMICALS**

In addition to the normal **fittings** and **reagents** in a CSE Chemistry laboratory, each candidate will require some of the listed apparatus and chemicals as will be prescribed in the 24 hours advance instructions.

**2.1 Apparati**

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|---|---|
| • 1 pipette (20 cm <sup>3</sup> or 25 cm <sup>3</sup> ) ✓ | • 1 measuring cylinder (100 cm <sup>3</sup> ) ✓ |
| • 1 burette (50 cm <sup>3</sup> ) ✓                       | • 1 measuring cylinder (10 cm <sup>3</sup> ) ✓  |
| • 3 titration flasks ✓                                    | • 1 stopwatch/clock ✓                           |
| • 1 beaker (100 cm <sup>3</sup> ) ✓                       | • 1 heat source/burner ✓                        |
| • 2 beakers (250 or 300 cm <sup>3</sup> ) ✓               | • 1 test tube holder ✓                          |
| • 1 spatula ✓   | • 1 test tube rack ✓                            |
| • stirrer ✓   |   |
| • 4 test tubes (pyrex) ✓                                  | • 1 Qualitative Analysis Guidance sheet ✓       |
| • 1 Petri dish/watch glass ✓                              | • 1 thermometer (0 – 100 °C). ✓                 |
|   | • 1 sheet (A4) white plain paper ✓              |

**2.2 Chemicals**

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|---|--|
| ✓• 500 cm <sup>3</sup> of distilled water                                     | ✓• 2 g sulphuric acid (H <sub>2</sub> SO <sub>4</sub> ) ✓            |
| ✓• 3 g sodium hydroxide (NaOH) ✓  | ✓• 3 g calcium carbonate (CaCO <sub>3</sub> ) ✓                      |
| ✓• 3 g sodium chloride (NaCl)   | ✓• 2 g hydrochloric acid (HCl) ✓                                     |
| ✓• 3 g sodium thiosulphate (Na <sub>2</sub> S <sub>2</sub> O <sub>3</sub> ) ✓ | ✓• 3 g anhydrous sodium carbonate (Na <sub>2</sub> CO <sub>3</sub> ) |
| ✓• 10 cm <sup>3</sup> methyl orange indicator (MO). ✓                         | ✓• 2 g nitric acid (HNO <sub>3</sub> ) ✓                             |
| ✓• 10 cm <sup>3</sup> phenolphthalein indicator (POP) ✓                       | ✓• 2 g oxalic acid ((COOH) <sub>2</sub> ·2H <sub>2</sub> O) ✓        |
| ✓• 3 g lead nitrate (Pb(NO <sub>3</sub> ) <sub>2</sub> ) ✓                    | ✓• 3 g potassium hydroxide (KOH) ✓                                   |
| ✓• 3 g copper carbonate (CuCO <sub>3</sub> )                                  | ✓• 3 g zinc sulphate (ZnSO <sub>4</sub> ·7H <sub>2</sub> O)          |
| ✓• 3 g calcium chloride (CaCl <sub>2</sub> ·6H <sub>2</sub> O)                | ✓• 3 g copper(II) sulphate (CuSO <sub>4</sub> ·5H <sub>2</sub> O) ✓  |
| ✓• 3 g ammonium chloride (NH <sub>4</sub> Cl)                                 | ✓• 3 g lead chloride (PbCl <sub>2</sub> ) ✓                          |
| ✓• 3 g ammonium hydroxide (NH <sub>4</sub> OH)                                |  |

- ✓• 3 g sodium hydrogen carbonate ( $\text{NaHCO}_3$ ).
- ✓• 3 g potassium ferrocyanide
- ✓• 3 g silver nitrate ( $\text{AgNO}_3$ ) ✓
- ✓• 3 g magnesium sulphate ( $\text{MgSO}_4$ )
- ✓• 3 g barium chloride ( $\text{BaCl}_2$ )
- ✓• 3 g potassium dichromate.